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Section 7.5 The Quantum Mechanical Model of the Atom

Section 7.7 Orbital Shapes and Energies

Section 7.11a How to Draw Orbital Diagrams for Elements

Section 7.11b How to Write a Complete Electron Configuration for an Element

Section 7.11c How to Write an Abbreviated Electron Configuration for an Element

Section 7.11d Electron Configurations for Cations and Anions

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Physical chemistry - Physical chemistry by Academic Lesson 364,541 views 4 years ago 11 hours, 59 minutes - Physical **chemistry**, is the study of macroscopic, and particulate phenomena in **chemical**, systems in terms of the principles, ...

Course Introduction

Concentrations

Properties of gases introduction

The ideal gas law

Ideal gas (continue)

Dalton's Law

Real gases

Gas law examples

Internal energy

Expansion work

Heat

First law of thermodynamics

Enthalpy introduction

Difference between H and U

Heat capacity at constant pressure

Hess' law

Hess' law application

Kirchhoff's law

Adiabatic behaviour

Adiabatic expansion work

Heat engines

Total carnot work

Heat engine efficiency

Microstates and macrostates

Partition function

Partition function examples

Calculating U from partition

Entropy

Change in entropy example

Residual entropies and the third law

Absolute entropy and Spontaneity

Free energies

The gibbs free energy

Phase Diagrams

Building phase diagrams

The clapeyron equation

The clapeyron equation examples

The clausius Clapeyron equation

Chemical potential

The mixing of gases

Raoult's law

Real solution

Dilute solution

Colligative properties

Fractional distillation

Freezing point depression

Osmosis

Chemical potential and equilibrium

The equilibrium constant

Equilibrium concentrations

Le chatelier and temperature

Le chatelier and pressure

Ions in solution

Debye-Huckel law

Salting in and salting out

Salting in example

Salting out example

Acid equilibrium review

Real acid equilibrium

The pH of real acid solutions

Buffers

Rate law expressions

2nd order type 2 integrated rate

2nd order type 2 (continue)

Strategies to determine order

Half life

The arrhenius Equation

The Arrhenius equation example

The approach to equilibrium

The approach to equilibrium (continue..)

Link between K and rate constants

Equilibrium shift setup

Time constant, tau

Quantifying tau and concentrations

Consecutive chemical reaction

Multi step integrated Rate laws

Multi-step integrated rate laws (continue..)

Intermediate max and rate det step

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Solutions, In this video Paul Andersen explains the important properties of **solutions**,. A **solution**, can be either a solid, liquid or ...

Solutions

Separation

Column Chromatography

Distillation

Formation of Solution

moles of solute

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Balance the Number of Oxygen Atoms

Single Replacement Reactions

Aluminum Reacting with Nickel to Chloride

Zinc Metal Reacting with Hydrochloric Acid

Silver Nitrate Reacting with Magnesium Fluoride

Precipitation Reaction

Sodium Carbonate with Hydrochloric Acid

Gas Evolution Reaction

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Intro

Number 10 Squares

Number 9 Diagrams

Number 8 Picture

Number 7 Picture

Number 6 Picture

Number 5 Picture

Number 4 Picture

Number 3 Elephant

Number 2 Squares

Solutions Overview and Types - Solutions Overview and Types by Tyler DeWitt 261,698 views 3 years ago 12 minutes, 16 seconds - This is an overview of **solutions**, or homogeneous mixtures, which have a uniform and even composition. They are different from ...

Introduction

Solutions vs Not Solutions

Parts

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